PAPER SOLUTION

From Meerut

JEE MAIN 2026

JAN	SHIFT
23	2nd

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#Q. In estimation of chlorine by Carius method, 0.245 g organic compound gave 0.5453 g AgCl. Find percentage of chlorine in the organic compound

Ans. (55)

#Q. How many of the following complex(es) have unpaired electrons
$[Ni(CO)_4]$, $[NiCl_4]^{2-}$, $[PtCl_4]^{2-}$, $[Pt(CN)_4]^{2-}$

Ans. (1)

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#Q. For XeO_2F_2, select the correct statement(s).

(A) It shows see-saw shape.

(B) Number of lone pair(s) Of e^- on Xe is 1.

(C) $\angle FXeF = 180^o$ (approx.)

(D) It has tetrahedral shape.

A (A), (C), (D) only

B (A), (B) only

C (A), (B), (C) only

D (B), (C), (D) only

Ans. (C)

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#Q. Given below are two statements

Statement I : Size of O^{2-} is smaller than F^-.

Statement II : Electronegativity of F is more than that of oxygen.

In the light of above statements, choose the correct option.

A. Statement I and statement II both are correct

B. Statement I and statement II both are incorrect

C. Statement I correct statement II incorrect

D. Statement I incorrect statement II correct

Ans. (C)

#Q. An ideal solution is formed by mixing 3 mole of A and 1 mole of B and the vapour pressure of solution is found to be 500 mm Hg. After further addition of 1 mole A, pressure of solution becomes 520 mm Hg. Find P_A^o.

Ans. (600)

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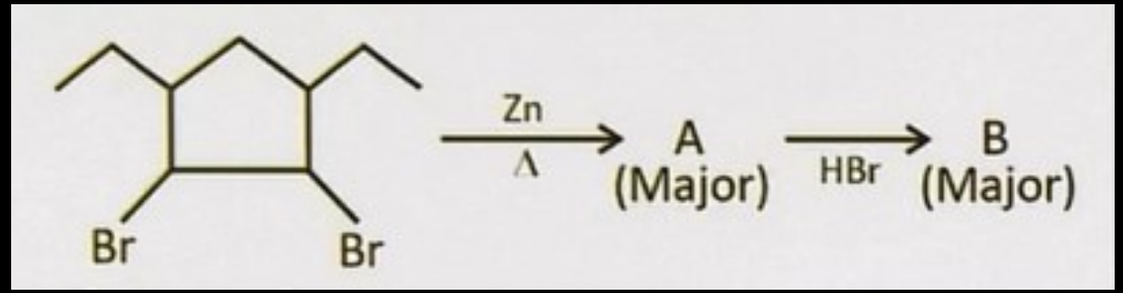
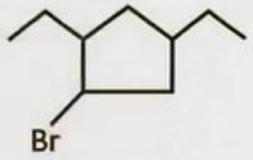
#Q. Consider the reaction,



Choose the correct option,

A A is 

B B is 

C B is 

D A is 

Ans. (C)

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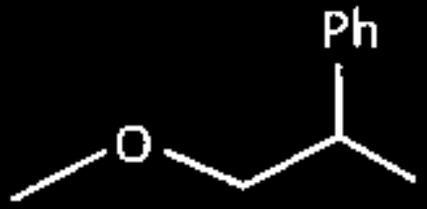
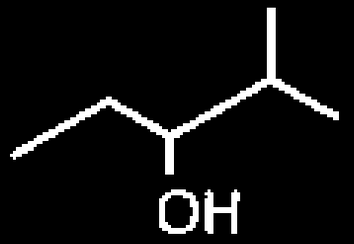
#Q. Which of the following molecule gives iodoform reaction.

A 

B 

C CH_3CH_2OH

D CH_3COOH

Ans. (C)

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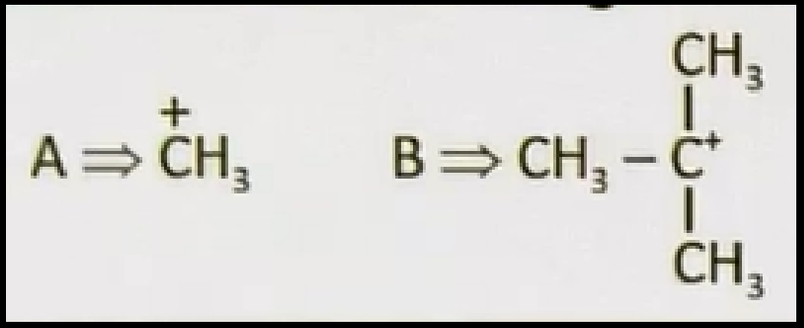
#Q. Consider the following intermediates.



A B is more stable than A as it has 9 α hydrogen

B A is more stable than B as it has 3 α hydrogen

C B is more stable than A due to resonance

D A is more stable due to inductive effect

Ans. (A)

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#Q. What is the oxidation state of chromium in the product when $K_2Cr_2O_7$ reacts with acidified KI

A. +6

B. +3

C. +4

D. +5

Ans. (B)

#Q. 250 cc of $x \times 10^{-3}$ M acidified $K_2Cr_2O_7$ solution neutralises 750 cc of 0.6 M Mohr's salt. Value of x is

(A) 200

(B) 600

(C) 400

(D) 300

Ans. (D)

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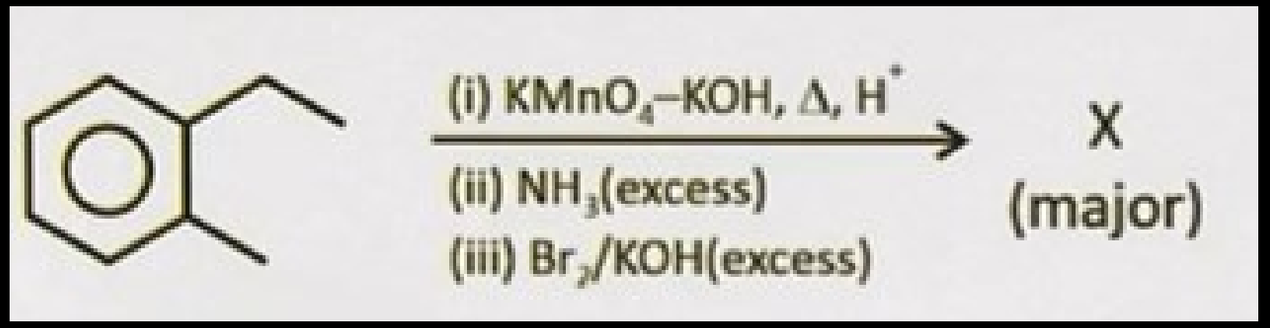
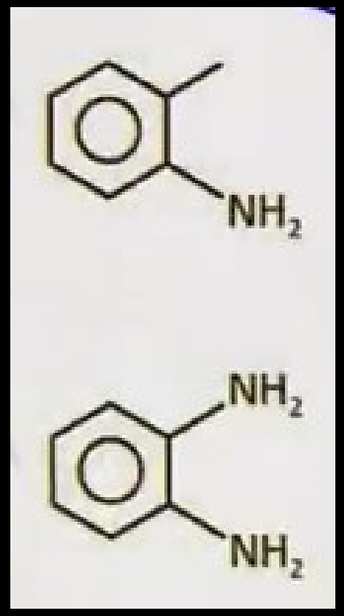
#Q. Consider the reaction,

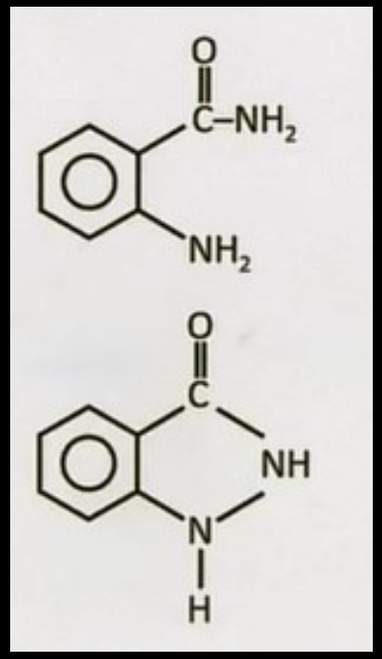


X is,

A



B



C

D

Ans. (C)

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#Q. Consider the following statement about complexes and its hybridisation.

A. $[CoF_6]^{3-}$; outer orbital complex, sp^3d^2

B. $[Ni(CN)_4]^{2-}$; inner orbital complex, dsp^2

C. $[Co(NH_3)_6]^{3+}$; inner orbital complex; d^2sp^3

D. $[FeF_6]^{3-}$; outer orbital complex; sp^3d^2

Choose the correct statement.

(A) A, B, cand D

(B) A, B and C only

(C) A and B only

(D) B and C only

Ans. (A)

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#Q. Consider the following reactions

$PbO_2 + Pb \rightleftharpoons 2PbO$; ΔG_1

$SnO_2 + Sn \rightleftharpoons 2SnO$; ΔG_2

Select the correct option?

A $\Delta G_1 > 0$, $\Delta G_2 > 0$

B $\Delta G_1 < 0$, $\Delta G_2 > 0$

C $\Delta G_1 < 0$, $\Delta G_2 < 0$

D $\Delta G_1 > 0$, $\Delta G_2 < 0$

Ans. (B)

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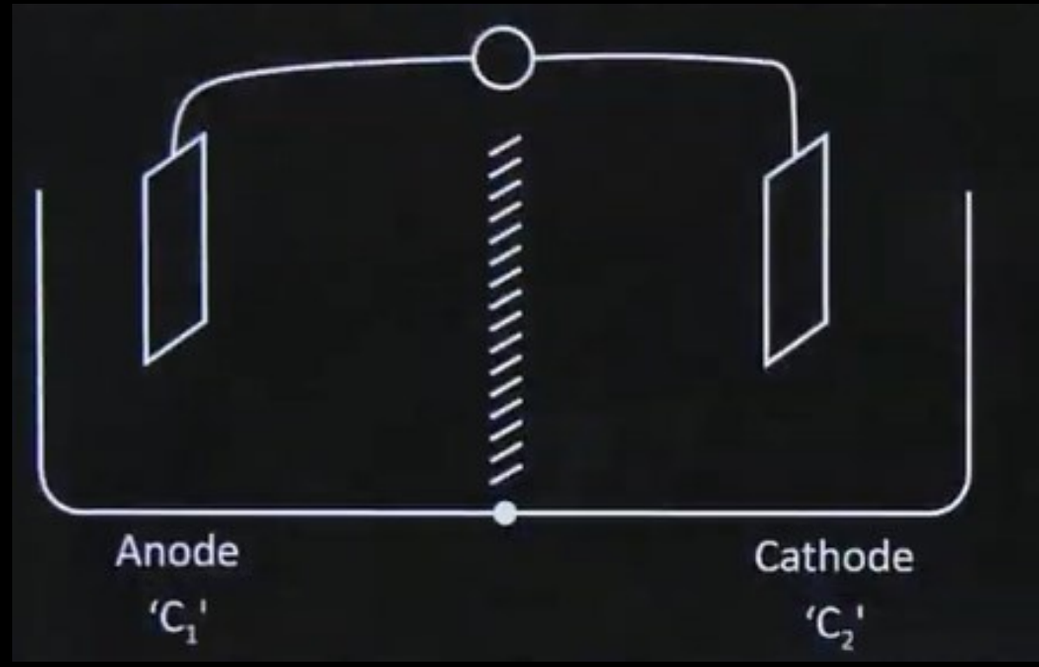
#Q. Consider a galvanic cell, made up of two Hrelectrodes,



Both compartments contain the same metal electrodes. If concentrations of in anode and cathode are C_1 and C_2 respectively, then $E_{cell} > 0$ when, (Pressure of H_2 = 1 atm)

A $C_2 < C_1$

B $C_2 = C_1$

C $C_2 > C_1$

D $C_2 < 0.5\ C_1$

Ans. (C)

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#Q. A compound P with molecular formula C_6H_7N is sparingly soluble in water. However on reaction with HCl. It becomes soluble. On reaction with KOH + $CHCl_3$, it gives foul smelling compound Q. The number of different type of H atoms present in P is:

(A) 4

(B) 5

(C) 7

(D) 8

Ans. (A)

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#Q. DNA is optically active due to the presence of:

(A) Purine nitrogenous base

(B) Phosphate molecule

(C) D-pentose sugar

(D) L-pentose sugar

Ans. (C)

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#Q. Two metals with work function in ratio 1 : 2, are exposed with photons of energy 6 eV. If $KE_A : KE_B$ is 2.642 : 1, then $\emptyset_A$ and $\emptyset_B$ value (in eV) are:

(A) 2.3, 4.6

(B) 1.4, 2.8

(C) 2.3, 3.6

(D) 3.2, 6.4

Ans. (A)